aqueous NH_4Cl (10 mL) was poured into the reaction mixture. The aqueous layer was extracted with CH_2Cl_2 (3 × 15 mL). The organic phases were dried over $MgSO_4$. After evaporation of the solvent the crude material was purified by flash chromatography over silica gel (hexane/ethyl acetate/ethanol (6:3:1)) affording 1.08 g (42%) of **8** and 0.46 g (18%) of **9**. Compound **8**: oil; $[\alpha]^{20}_D$ +18° (*c* 2.55, $CHCl_3$); IR (cm^{-1}, neat) 3550; ^1H NMR ($CDCl_3$) δ 2.72–2.82 (m, 1 H), 3.07–3.18 (m, 2 H), 3.15 (s, 3 H), 3.36 (s, 3 H), 3.77 (s, 3 H), 4.08–4.12 (m, 1 H), 4.28 and 4.48 (AB dd, 2 H, J = 6.8 Hz), 4.53 and 4.57 (AB dd, 2 H, J = 7.0 Hz), 6.83 and 7.18 (AB dd, 4 H, J = 8.6 Hz); MS m/e (M^+ − 18) 309, 280, 190, 121, 96, 45. Anal. Calcd for $C_{16}H_{25}NO_6$ (327.37): C, 58.69; H, 7.69; N, 4.27. Found: C, 58.57; H, 7.61; N, 4.32.

(2*R*,3*S*,4*S*)-2-(4-Methoxybenzyl)-3,4-bis(methoxymethoxy)pyrrolidine (10). Compound **8** (0.5 g, 1.52 mmol) was dissolved in MeOH (20 mL) and hydrogenated at 1 atm in the presence of Raney Ni (0.07 g) for 6 h at rt. The catalyst was removed by filtration and washed thoroughly with methanol. After evaporation of the solvent, the crude product was purified by flash chromatography (hexane/ethyl acetate/ethanol/30% NH_4OH (5:3:1.8:0.2)) to give 0.4 g (85%) of amine **10** as an oil: $[\alpha]^{20}_D$ −11.6° (*c* 2.35, MeOH) (lit.[7c] $[\alpha]^{20}_D$ −12.5° (*c* 1.52, MeOH); IR (cm^{-1}, neat) 3300; ^1H NMR ($CDCl_3$) δ 2.02 (bs, 1 H), 2.79 (dd, 1 H, J = 8.0, 17.7 Hz), 2.84 (dd, 1 H, J = 6.2, 13.7 Hz), 2.92 (m, 1 H), 3.29 (s, 3 H), 3.32–3.42 (m, 2 H), 3.39 (s, 3 H), 3.78 (s, 3 H), 3.82 (dd, 1 H, J = 4.7, 1.5 Hz), 4.12 (dt, 1 H, J = 3.4, 1.5 Hz), 4.50 and 4.64 (AB dd, 2 H, J = 6.7 Hz), 4.63 and 4.69 (AB dd, 2 H, J = 6.7 Hz), 6.83 and 7.17 (AB dd, 4 H, J = 8.7 Hz); MS m/e M^+ 312, 304, 190, 121, 96, 68, 45.

(2*R*,3*S*,4*S*)-2-(4-Methoxybenzyl)pyrrolidine-3,4-diol (Deacetylanisomycin, 11). Pyrrolidine **10** (0.15 g, 0.48 mmol) was refluxed in a mixture of MeOH/6 N HCl (1:1) for 24 h. The solution was then concentrated, made alkaline with Na_2CO_3, and extracted with $CHCl_3$ (2 × 30 mL). The aqueous phase was allowed to stand overnight at 5 °C, and the colorless crystals were collected by suction giving 0.08 g (75%) of **11**: mp 171–172 °C (lit.[7c] mp 176–177 °C).

Acknowledgment. The authors wish to thank the Ministero dell'Università e della Ricerca Scientifica e Tecnologica of Italy for financial assistance.

Registry No. **1**, 22862-76-6; **4**, 90365-74-5; **5**, 137945-72-3; **6**, 138051-80-6; **7**, 137945-73-4; **8**, 137945-74-5; **9**, 137945-75-6; **10**, 100449-58-9; **11**, 27958-06-1; (4-methoxybenzyl)magnesium bromide, 38769-92-5.

Additions and Corrections

Vol. 45, 1980

Arthur G. Anderson, Jr.,* Gary M. Masada, and Glenn L. Kao. Electrophilic Trifluoroacetylation of Dicyclopenta[*ef,kl*]-heptalene (Azupyrene).

Page 1313, left column, line 2, should read the 1 position and the latter the 4 position.

Vol. 52, 1987

Arthur G. Anderson, Jr.,* and Edward D. Daugs. Dicyclopenta[*ef,kl*]heptalene (Azupyrene) Chemistry. Electrophilic Monosubstitution: Acetylation, Halogenation, and Thiocyanation. 1-(Ethoxymethyl)azupyrene and Dimethyl (1-Azupyrenylmethyl)malonate. Acetylazupyrene Geometry..

Page 4392, left column. The reference to the ^1H NMR signal for H-2 of 1-(trifluoroacetyl)azupyrene (**10**) should be 3a instead of 7. The references to the ^1H–^{19}F coupling in the analogous 1-(trifluoroacetyl)azulene and phenyl hexafluorobutyl ketone should be 7 and 8, respectively.

Vol. 56, 1991

Arthur G. Anderson, Jr.* and Ralph D. Haddock. The Thermal Isomerization of [2a,11-$^{13}C_2$]Dicyclopenta[*ef,kl*]heptalene (Azupyrene) to Pyrene.

Page 551, right column, Scheme III, above the left-hand arrow should read "$\pi_s^2 + \pi_a^2$". Line 5 under Scheme III should read "also 1- and 2-methylpyrene (**4** and **5**) from 1-methylazupyrene".

Suruliappa Jeganathan and Pierre Vogel*. Highly Stereoselective Total Syntheses of Octoses and Derivatives.

Page 1135. Reference 46 should have the following added: Bilik, V.; Petrus, L.; Aldölfi, J. *Chem. Zvesti* **1976**, *30*, 698.

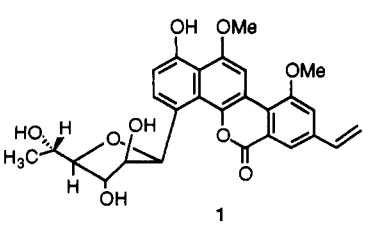
Page 1137, column 2, line 12, should read (−)-(**7**) were obtained as hydroscopic solid materials.

Page 1136, column 1, lines 11 and 13, page 1137, column 1, line 19 and 21, page 1140, column 2, lines 31–35, lines 41–44, and page 1141, column 2, lines 43–46 and lines 52–55: because of the presence of impurities due to incomplete hydrolysis of the acetonide precursors of octoses (−)-**4**, (+)-**5**, (+)-**6**, and (−)-**7**, the ^1H-NMR signal attributions for the anomeric protons of these carbohydrates as well as the proportions given for the corresponding α-furanose, β-furanose, α-pyranose, and β-pyranose forms cannot be considered as definitive.

We thank Professor S. J. Angyal, the University of New South Wales, Australia, for pointing out these problems to us.

Kathlyn A. Parker* and Craig A. Coburn. A Strategy for the Convergent Synthesis of Gilvocarcins via Chromium Carbene Benzannulation. 1-*O*-Methyldefucogilvocarcin V in Seven Steps.

Page 1666. Structure **1** should be



1

James M. Tanko,* N. Kamrudin Suleman, and Joseph F. Blackert. Kinetic vs Thermodynamic Factors in α-Hydrogen Atom Abstractions from Alkylaromatics. 2. Reactivities of α-Alkylnaphthalenes and Several Conformationally Locked Alkylaromatics toward Bromine Atom.

Page 6395. The author name N. Kamrudin should be N. Kamrudin Suleman.